TABLE XIX. Ionic Valences

POSITIVE IONS		NEGATIVE IONS	
Name	*Formula*	*Name*	*Formula*
Hydrogen	H^+	Acetate	CH_3COO^-, or $C_2H_3O_2^-$
Ammonium	NH_4^+		
Lithium	Li^+	Fluoride	F^-
Potassium	K^+	Chloride	Cl^-
Silver	Ag^+	Bromide	Br^-
Sodium	Na^+	Iodide	I^-
Copper I (cuprous)	Cu^+	Hypochlorite	ClO^-
Copper II (cupric)	Cu^{++}	Chlorite	ClO_2^-
Mercury I (mercurous)	Hg_2^{++}	Chlorate	ClO_3^-
Mercury II (mercuric)	Hg^{++}	Perchlorate	ClO_4^-
Barium	Ba^{++}	Nitrate	NO_3^-
Cadmium	Cd^{++}	Nitrite	NO_2^-
Calcium	Ca^{++}	Carbonate	CO_3^{--}
Lead (plumbous)	Pb^{++}	Hydrogen carbonate (bicarbonate)	HCO_3^-
Magnesium	Mg^{++}	Chromate	CrO_4^{--}
Strontium	Sr^{++}	Dichromate (bichromate)	$Cr_2O_7^{--}$
Zinc	Zn^{++}	Oxalate	$C_2O_4^{--}$
Cobalt II (cobaltous)	Co^{++}	Sulfate	SO_4^{--}
Cobalt III (cobaltic)	Co^{+++}	Hydrogen sulfate (bisulfate)	HSO_4^-
Nickel II (nickelous)	Ni^{++}	Sulfite	SO_3^{--}
Nickel III (nickelic)	Ni^{+++}	Hydrogen sulfite (bisulfite)	HSO_3^-
Manganese II (manganous)	Mn^{++}	Sulfide	S^{--}
Manganese III (manganic)	Mn^{+++}	Hydrogen sulfide ion (bisulfide ion)	HS^-
Iron II (ferrous)	Fe^{++}	Thiosulfate	$S_2O_3^{--}$
Iron III (ferric)	Fe^{+++}	Peroxydisulfate	$S_2O_8^{--}$
Tin II (stannous)	Sn^{++}	Borate	BO_3^{---}
Tin IV (stannic)	Sn^{++++}	Phosphate (ortho)	PO_4^{---}
Aluminum	Al^{+++}	Dihydrogen phosphate	$H_2PO_4^-$
Bismuth	Bi^{+++}	Monohydrogen phosphate	HPO_4^{--}
Scandium	Sc^{+++}	Metaphosphate	PO_3^-
Chromium II (chromous)	Cr^{++}	Phosphite	HPO_3^{--}
Chromium III (chromic)	Cr^{+++}	Ferrocyanide	$Fe(CN)_6^{----}$
Hydroxide	OH^-	Ferricyanide	$Fe(CN)_6^{---}$