**Table 16.2 Standard electrode potentials at 25° C[a]**

| HALF REACTION | $\mathcal{E}^\circ$ (VOLTS) |
|---|---|
| $Li^+ + e \rightleftharpoons Li$ | −3.045 |
| $K^+ + e \rightleftharpoons K$ | −2.925 |
| $Ba^{2+} + 2e \rightleftharpoons Ba$ | −2.906 |
| $Ca^{2+} + 2e \rightleftharpoons Ca$ | −2.866 |
| $Na^+ + e \rightleftharpoons Na$ | −2.714 |
| $Mg^{2+} + 2e \rightleftharpoons Mg$ | −2.363 |
| $Al^{3+} + 3e \rightleftharpoons Al$ | −1.662 |
| $2H_2O + 2e^- \rightleftharpoons H_2 + 2OH^-$ | −0.82806 |
| $Zn^{2+} + 2e \rightleftharpoons Zn$ | −0.7628 |
| $Cr^{3+} + 3e \rightleftharpoons Cr$ | −0.744 |
| $Fe^{2+} + 2e \rightleftharpoons Fe$ | −0.4402 |
| $Cd^{2+} + 2e \rightleftharpoons Cd$ | −0.4029 |
| $Ni^{2+} + 2e \rightleftharpoons Ni$ | −0.250 |
| $Sn^{2+} + 2e \rightleftharpoons Sn$ | −0.136 |
| $Pb^{2+} + 2e \rightleftharpoons Pb$ | −0.126 |
| $2H^+ + 2e \rightleftharpoons H_2$ | 0 |
| $Cu^{2+} + 2e \rightleftharpoons Cu$ | +0.337 |
| $Cu^+ + e \rightleftharpoons Cu$ | +0.521 |
| $I_2 + 2e \rightleftharpoons 2I^-$ | +0.5355 |
| $Fe^{3+} + e \rightleftharpoons Fe^{2+}$ | +0.771 |
| $Ag^+ + e \rightleftharpoons Ag$ | +0.7991 |
| $Br_2 + 2e \rightleftharpoons 2Br^-$ | +1.0652 |
| $O_2 + 4H^+ + 4e \rightleftharpoons 2H_2O$ | +1.229 |
| $Cr_2O_7^{2-} + 14H^+ + 6e \rightleftharpoons 2Cr^{3+} + 7H_2O$ | +1.33 |
| $Cl_2 + 2e^- \rightleftharpoons 2Cl^-$ | +1.3595 |
| $MnO_4^- + 8H^+ + 5e \rightleftharpoons Mn^{2+} + 4H_2O$ | +1.51 |
| $F_2 + 2e \rightleftharpoons 2F^-$ | +2.87 |

[a] Data from A. J. de Bethune and N. A. Swendeman Loud, "Table of Electrode Potentials and Temperature Coefficients," pp. 414–424 in *Encyclopedia of Electrochemistry* (C. A. Hampel, editor), Van Nostrand Reinhold, New York, 1964, and from A. J. de Bethune and N. A. Swendeman Loud, *Standard Aqueous Electrode Potentials and Temperature Coefficients*, 19 pp., C. A. Hampel, publisher, Skokie, Illinois, 1964.